	SO_4^{2-}	SO_3^{2-}	$S_2O_3^{2-}$	PO_4^{3-}	CO_3^{2-}	CrO_4^{2-}	F^-	$C_2O_4^{2-}$
$AgNO_3$	No precipitate	Ag_2SO_3 white precipitate, soluble in excess SO_3^{2-}, soluble in HNO_3, CH_3COOH, and HN_3 aq Ag_2SO_3 heated with water decomposes into metallic silver and H_2SO_4	Ag_2S_2O3 white fast yellowing and then blackening precipitate, soluble in excess $S_2O_3^{2--}$	Ag_3PO_4 yellow precipitate	Ag_2CO_3 white precipitate, after heating Ag_2O brown precipitate	$Ag_2Cr_2O_7$ brown-red precipitate	No precipitate	Ag_2CrO_4 white precipitate, soluble in dilute HNO_3 and NH_{3aq}
$Ba(NO_3)_2$	$BaSO_4$ white precipitate (insoluble in excess dilute strong acids)	$BaSO_3$ white precipitate (soluble in excess dilute strong acids)	BaS_2O_3 white precipitate, soluble in hot water and dilute HNO_3 and HCl	$BaHPO_4$ white precipitate	$BaCO_3$ white precipitate	$BaCrO_4$ yellow precipitate	BaF_2 white precipitate	BaC_2O_4 white precipitate, soluble in dilute HNO_3 and NH_3aq
$Pb(NO_3)_2$	$PbSO_4$ white precipitate			$Pb_3(PO_4)_2$ white precipitate	$PbCO_3$ white precipitate	$PbCrO_4$ yellow precipitate	PbF_2 white precipitate	
		SO_2 liberated on warming with dilute acids			CO_2 liberated by dilute acids (HCl)		$Fe(SCN)^{2+}$+ 6F- + H^+= FeF_6^{3-}+ $2SCN^-$ discoloration of the solution	CaC_2O_4 white precipitate, soluble in mineral acids and CH_3COOH

Table 1. Characteristic reactions of anions of Group I